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waves in a vacuum is the same, waves can have different wavelengths and frequencies. As you can see from the equation on the previous page, wavelength and frequency are inversely related; in other words, as one quantity increases, the other decreases.

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Chapter 5 – Electrons in Atoms. Chapter 5 – Electrons in Atoms. Jennie L. Borders. Section 5.1 – Models of the Atom. The Rutherford's model of the atom did not explain how an atom can emit light or the chemical properties of an atom. Plum Pudding Model Rutherford's Model. The Bohr Model.

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Directions: Mark the best possible answer choice for each test question below on your Scan-Tron card. Use the diagram to answer questions 1-4. 1) The visible

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light is being \_\_\_\_\_ through a prism to produce a spectrum of colors.

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Chapter 5 Electrons in Atoms 2. Light and Quantized Energy (5.1) 

- The study of light led to the development of the quantum mechanical model.

- Light is a kind of electromagnetic radiation (EM).

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All move at  $3.00 \times 10^8$  m/s (c) Speed of light. 3.

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116 Chapter 5 Electrons in Atoms CHAPTER 5 What You'll Learn You will compare the wave and particle models of light. You will describe how the frequency of light emitted by an atom is a unique characteristic of that atom. You will compare and contrast the Bohr and quantum mechanical models of the atom.

[Unit 4: Electrons in Atoms \(Chapter 5\) You'll Remember ...](#)

Chapter 5: Electrons in Atoms. Models of the Atom Rutherford used existing ideas about the atom and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun. Rutherford's model fails to explain why objects change color when heated.

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Chapter 5: Electrons in Atoms Models of the Atom • Rutherford used existing ideas about the atom and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun.

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Unformatted text preview: Chapter 5 “Electrons in Atoms” Section 5.1 Models of the Atom • OBJECTIVES: •Identify the inadequacies in the Rutherford atomic model. Section 5.1 Models of the Atom • OBJECTIVES: •Identify the new proposal in the Bohr model of the atom. Section 5.1 Models of the Atom • OBJECTIVES: •Describe the energies and positions of electrons according to the ...

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Chapter 5 “Electrons in Atoms” ... Electrons would surround and move around it, like planets around the sun; Atom is mostly empty space; It did not explain the

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chemical properties of the elements – a better description of the electron behavior was needed; ... The Math in Chapter 5.

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Chemistry - Chp 5 - Electrons In Atoms - Powerpoint 1. Chapter 5“Electrons in Atoms” 2. Section 5.1 Models of the AtomOBJECTIVES:• Identify the inadequacies in the Rutherford atomic model. 3. Section 5.1 Models of the AtomOBJECTIVES:• Identify the new proposal in the Bohr model of the atom.

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4.

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Chapter 5 Electrons in Atoms . Name Date 11. The number of sublevels in an energy level is equal to the square of the principal quantum number of that energy level. 12. The maximum number of electrons that can occupy the fourth principal energy level of an atom is 32. 13. The higher the energy level occupied by an electron the more

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116 Chapter 5 Electrons in Atoms CHAPTER 5 What You'll Learn You will compare the wave and particle models. Key Concepts. withChemASAP. 5.1 Models of ... 5.3 Physics and the Quantum Mechanical Model 3 d. n. 4. 5.2 Electron Arrangement in Atoms....

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CHAPTER 5 ELECTRONS IN ATOMS. How do electrons behave in atoms ? Rutherford proposed that electrons move around the nucleus like planets around the sun. Rutherford's atomic model could not explain the chemical properties of elements. THE BOHR MODEL: Niels Bohr – 1885 – 1962.

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116 Chapter 5 Electrons in Atoms CHAPTER 5 What You'll Learn You will compare the wave and particle models of light. You will describe how the frequency of light emitted by an atom is a unique characteristic of that atom.

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Chapter 5: Electrons in Atoms Chapter 5: Electrons in Atoms CHEMISTRY  
Matter and Change Section 5.1 Light and Quantized Energy Section 5.2 Quantum



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Theory and the Atom Section 5.3 Electron Configuration Exit Table Of Contents  
CHAPTER 5 Click a hyperlink to view the corresponding slides.

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Now that the 2s subshell is filled, electrons in larger atoms must go into the 2p subshell, which can hold a maximum of six electrons. The next six elements progressively fill up the 2p subshell: B:  $1s^2 2s^2 2p^1$  C:  $1s^2 2s^2 2p^2$  N:  $1s^2 2s^2 2p^3$  O:  $1s^2 2s^2 2p^4$  F:  $1s^2 2s^2 2p^5$  Ne:  $1s^2 2s^2 2p^6$

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CHAPTER Date' Class 5.2 STUDY GUIDE FOR CONTENT MASTERY hertz speed Electrons in Atoms Section 5.1 Light and Quantized Energy In your textbook, read about the wave nature of light. Use each of the terms below just once to complete the passage. amplitude light energy wave frequency wavelength

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